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~~CHEM 131 L Experiment 9 The questions and answers for ...~~

~~Answers to Post-Lab Questions for Experiment~~

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5. I have 40 L of methane gas measured at standard temperature and pressure (STP). How many moles of methane gas is this? 1 mole of any gas at STP = 22.4 L Moles of methane = $(40 / 22.4) = 1.79$ moles. What is the volume of 0.30 moles of nitrogen gas if measured at STP? 1 mole of any gas at STP = 22.4 L Volume of nitrogen = $0.3 \times 22.4 = 6.72$ L.

~~SCC1123 Experiment 5 Post Lab Questions Answers — ECU ...~~

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Favorite Answer. First calculate the moles of KHP (potassium hydrogen phthalate, $\text{KHC}_8\text{H}_4\text{O}_4$). The molar mass for KHP = 204.2 g/mole. $0.531 \text{ g KHP} \times (1 \text{ mole KHP} / 204.2 \text{ g KHP}) = 0.00260$ moles KHP. KHP...

~~Chemistry post lab question? | Yahoo Answers~~

Favorite Answer. The white precipitate most likely came from the presence of chloride ions. $\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl}(s)$ The chemical responsible for the iron is Fe (iron), in some form. It could be almost...

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Follow this series of questions. When you can answer " yes: " to a question, then stop. 1) Does your reaction have oxygen as one of it's reactants and carbon dioxide and water as products? Yes: it's a combustion reaction 2) Does your reaction have two (or more) chemicals combining to form one chemical? Yes: it's a synthesis reaction

~~Chemistry of Copper~~

Sear Chemistry 200 - Organic Chemistry - Post-Lab Assignment - Extraction - F2020 (2) - Saved Layout References Mailings Review View Help Acrobat TL Design Name This assignment will guide you through an extraction procedure. Your task is to separate a mixture of 3 compounds using a series of acid-base extractions.

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~~Chemistry 2A Lab Manual~~

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