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Chemistry How Sublime Laying The Sublimation is the transition from the solid phase to the gas phase without passing through an intermediate liquid phase. This endothermic phase transition occurs at temperatures and pressures below the triple point. The term only applies to physical

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Sublimation is the transition of a substance directly from the solid to the gas state, without passing through the liquid state. Sublimation is an endothermic process that occurs at temperatures and pressures below a substance's triple point in its phase diagram, which corresponds to the lowest pressure at which the substance can exist as a liquid. The reverse process of sublimation is deposition or desublimation, in which a substance passes directly from a gas to a solid phase. Sublimation has

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Sublimation in chemistry refers to the phase transition in which matter changes state from a solid immediately into a gas, without passing through an intermediate liquid phase. Sublimation occurs when atmospheric pressure is too low for a substance to exist in liquid form. Sublimation is the inverse of deposition, the phase transition in which gas goes immediately to a solid state.

What Is Sublimation In Chemistry? | Science Trends
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